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Comments. As a result, more  $\text{Cu}(\text{OH})_2$  should precipitate from the solution. This is the common ion effect. 12 0

obj The amount of  $\text{NaCl}$  that could dissolve to reach the saturation point would be

lowered. Example: Let's determine what fraction of the total iron (III)

concentration is present as the  $\text{Fe}^{3+}$  ion in a solution that was initially 0.10 M

$\text{Fe}^{3+}$  and 1.0 M  $\text{SCN}^-$ ?

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[Cl<sup>-</sup>] = 0.10 M (due to ...

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pogil common ion effect on solubility answers. In the water treatment process, sodium carbonate salt is added to precipitate the calcium carbonate. The initial value of the reaction quotient is therefore equal to zero. Through the addition of common ions, the solubility of a compound generally decreases due to a shift in equilibrium.

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Use your solubility curve graph provided to answer the following questions. 1. What are the customary units of solubility on solubility curves? Yemp 2. Define solubility. ok be aya 3. According to the graph, the solubility of any substance changes as changes. 4. List the substances whose solubility decreases as temperature increases. 5.

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POGIL 14. The solubility  
product constant,  $K_{sp}$ , for  
silver carbonate,  $Ag_2CO_3$ ,  
is  $8.10 \times 10^{-12}$ . Compare the  
solubility of silver  
carbonate in water to its  
solubility in a 0.100 M  
solution of silver nitrate.

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and b consistent with the  
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Model 3 - Number of Protons

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and Attractive Force Force  
of Attraction (Newtons)  $0.10$   
 $\text{nm}$   $2.30 \times 10^{-8}$   $0.10 \text{ nm}$   $\times$   
 $10^{-8}$   $0.10 \text{ nm}$   $6.90 \times$

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## Answers

Curves 5. Use the temperature dependence of solubility to identify whether the substances in Table G are gases or solids. Make two lists below, one for gases and one for solids. 6. Suggest a reason why solubility decreases with increasing temperature for gaseous solutes but increases for solid solutes. Problems 1.

### *Interpreting Solubility Curves*

b. Is your answer in part a consistent with your answer to Question 7: The enthalpy of fusion given in 'Model 2 for each substance is a measure of the energy (in

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k)) needed to melt a mole of that substance. a. Which type of solid has the largest enthalpies of fusion.' b.

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