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Lewis Theory
VII: Practice
with Lewis
Structures
*How to
Draw Lewis
Structures and
Calculate Formal
Charge*

Lewis Dot
Structure

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Practice Lewis

Problems (with
answers and
explanation) How
to Determine the
Most Stable

Lewis Structure
Practice

Problems,

Examples,

Questions,

Summary ~~Unit 7:~~

~~More Lewis~~

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To Draw The
Lewis Structure
of NO₃⁻ (Nitrate
Ion) Lewis dot
structures:

SCN . Formal
charges

How to Draw
Lewis
Structures!
(Lewis Dot
Diagrams)

Lewis Dot
Structures

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Lewis structure

- double \u0026amp;

triple bonds

VSEPR Theory

Practice

Problems

Lewis Dot

Structures: Easy

trick! ~~Lewis~~

~~Structures,~~

~~Introduction,~~

~~Formal Charge,~~

~~Molecular~~

~~Geometry,~~

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~~Resonance, Polar
or Nonpolar~~

Valence

Electrons and

the Periodic

Table *Lewis*

Structures Made

Easy: Examples

and Tricks for

Drawing Lewis

Dot Diagrams of

Molecules VSEPR

Theory:

Introduction

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*Electron Dot
Diagram*

*Lewis Electron
Dot Diagram*

*Lewis
Dot Structures
for Polyatomic
Ions*

CLEAR

*SIMPLE
Lewis Electron
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Dot Structures

~~Part 1 How to
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example) Lewis
Dot Structure
for elements~~

Tutorial:

Weighted Average

Atomic Mass

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Structures

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Monc LcW s

Structures tr I

Name Date: Hour:

Introduction

Questions Draw

the Lewis

structures for

the following

molecules. Part

c is done for

you. a) H₂O

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(oxygen is
central) b) SO_2
c) N_2 It t-t (a
[Jrt @ $\text{N}=\text{N}$ @ @
f\ @ 2 In

question 1 above
you should see
three different
kinds of bonds
of bonds as best
you can. a)
single bond: one
es. iv (e⁻)
shave a b) * -1;

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double bond:

Structures

Monocyclic Lewis

Answers -
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Name: _____

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Information:

Structure of the

Atom Note the

following

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symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms:

1 1H 2 1H 3 1H

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The following
three diagrams
are carbon
atoms: Notice
the type of
notation used
for atoms: 12 13
14 6C 6 C 6C (6
...

24 ChemQuest 8

Name: -

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Answer Key

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included in the
Lewis Structure.

$$4 + 3(6) + 2 =$$

24 (carbon has

four and each
oxygen has six;

add two for
the -2 charge) 5

$$+ 3(1) = 8$$

(nitrogen has
five; each

hydrogen has

one) Step #2:

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Draw the Lewis
“skeleton
structure” with
the central
atoms and the
other atoms,
each connected
with a single
bond. Step #3:
Add six more
electron dots to
each atom except
the

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Steps for

Writing Lewis

Structures.

Example \(\PageI

ndex{2}\) 1.

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Determine the total number of valence electrons in the molecule or ion. Each hydrogen atom (group 1) has one valence electron, carbon (group 14) has 4 valence electrons, and oxygen (group 16) has 6

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valence
electrons, for a
total of $[(2)(1)$
 $+ 4 + 6] = 12$

valence
electrons. 2.

10.5: Writing
Lewis Structures
for Covalent
Compounds ...
included in the
Lewis Structure.
 $4 + 3(6) + 2 =$

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24 (carbon has four and each oxygen has six; add two for the -2 charge) $5 + 3(1) = 8$ (nitrogen has five; each hydrogen has one) Step #2: Draw the "skeleton structure" with the central

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atoms and the
other atoms,
each connected
with a single
bond.

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